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exploration of the significance of language and its profound affect our existence. Throughout this critique, we shall delve into the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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modeling and molecular

situations are integrated

throughout underpinning this

text is the

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terms in this set 91

compressibility a measure of

how much the volume of matter

decreases under pressure

because of the space between

particles in a gas why are gases

easily compressed straight line

paths until they collide with

other particles or the walls of

their of container

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behavior learn the history of

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page 413 how does an

automobile air bag protect the

crash dummy from being

broken as a result of impact the

gases used to inflate the airbag

are able to absorb a

considerable amount of energy

when they are compressed 2

what theory explains the

behavior of gases kinetic

theory 3

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behavior of gases before you

read after you read this section

you should be able to answer

these questions what affects

how a gas behaves what are

the gas laws what affects the

behavior of a gas gases behave

differently than solids or

liquids gas particles have a

large amount of space between

them

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same as the volume of its 101

300 pa a pressure of 101 3 kpa

is equal to the air inside a fully

pumped basketball has a

higher pressure than the air

outside because there is a

higher number of gas particles

per unit volume does the air

inside a fully pumped

basketball have a higher or

lower pressure than the air

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containing terms like gases in

earth s atmosphere exert on

everything according to the

theory the particles of a gas are

moving pressure is this amount

of exerted per unit of area and

more

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directions answer each

question on the lines provided

you must include the terms

below in your answer boyle s

law charles s law kinetic

molecular theory pressure 1

temperature pressure and

volume affect the behavior of

gases which variable is held

constant in the relationship

described by boyle s law 2

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web the amount of space that

matter fills measured in cubic

centimeters cm³ milliliters ml

liters l gas particles move and

fill the space available this

measure of gas is the same as

the measure of its container a

measure of the average energy

of random motion of the

particles of a substance

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 unlike solids or liquids the
 reverse is also true they are
 easily compressed or squeezed
 into a smaller volume
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 how much the volume of matter
 decreases under pressure
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 no definite shape or volume
 and they expand to fill their
 container particles move
 rapidly in all directions
 molecules are in constant
 motion and collide with each
 other and walls of the
 container low density because
 particles are far apart are
 compressible
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 away to neolithic denmark and
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 epochs separated by thousands
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